

# Solution Stoichiometry Molarity Worksheet

## Chapter 1 : Solution Stoichiometry Molarity Worksheet

Molarity =  $\frac{\text{moles of solute}}{\text{liters of solution}}$  the molarity of a solution is a ratio of the moles of solute per liters of solution. the units for molarity are written as mol/l or m. this measurement is used to perform stoichiometric calculations. the strategy used for solving these solution stoichiometry problems is to set up

Solution stoichiometry worksheet solve the following solutions stoichiometry problems:

1. how many grams of silver chromate will precipitate when 150. ml of 0.500 m silver nitrate are added to 100. ml of 0.400 m potassium chromate?  
 $2 \text{ AgNO}_3(\text{aq}) + \text{K}_2\text{CrO}_4(\text{aq}) \rightarrow \text{Ag}_2\text{CrO}_4(\text{s}) + 2 \text{ KNO}_3(\text{aq})$   
0.150 l  $\text{AgNO}_3$  0.500 moles  $\text{AgNO}_3$  1 moles  $\text{Ag}_2\text{CrO}_4$  331 Nsm12c\$pal\$worksheet\$\$ molarity\$&\$stoichiometry\$ name:\$\_\_\_\_\$ \$ 6. 0.2500\$g\$hydrofluoric\$acid\$addedto50.0ml\$of\$0.500\$M\$NaOH\$solution.\$
2. Chemistry: molarity and stoichiometry directions: using the definition of molarity, the given balanced equations, and stoichiometry, solve the following what mass of calcium hydroxide is needed to make 2.0 liters of a 2.0 m calcium hydroxide solution? c. how many grams of calcium chloride will be made when 1.0 liter of a 1.0 m calcium chloride solution is titrated with 24.5 ml of 0.250 m sodium hydroxide solution?  
 $\text{CaCl}_2 + \text{NaOH} \rightarrow \text{Ca(OH)}_2 + \text{NaCl}$
3. titrations, etc practice problems! page 3 1. 13.0 ml of 0.250 m EDTA solution is titrated with 24.5 ml of 0.100 m  $\text{Mg}^{2+}$  solution. calculate the concentration of the EDTA solution.  
 $\text{Mg}^{2+} + \text{EDTA} \rightarrow \text{MgEDTA}^{2-}$

Chapter 4: worksheet #4 stoichiometry in aqueous reactions

1. if 30.0 ml of 0.150 m aluminum chloride is added to 15.0 ml of 0.100 m silver(i) nitrate, what mass of precipitate will be formed?
2. a 10.00 ml sample of barium hydroxide solution was titrated with 31.24 ml of 0.1251 m HBr. calculate the molarity of the barium hydroxide solution.

3. Solution stoichiometry in solids, moles are obtained by dividing mass by the molar mass in liquids, it is necessary to convert volume into moles using molarity

Molarity worksheet #1 for each of the following problems, use proper units and show all work:

1. if 10.7 grams of  $\text{NH}_4\text{Cl}$  is dissolved in enough water to make 800 ml of solution, what will be its molarity? (answer: 0.25 mol/l).
2. calculate the molarity of a solution prepared by dissolving 6.80 grams of  $\text{AgNO}_3$  in enough water to make 250.0 ml of solution.

Homework answers molarity & molality worksheet student due date 3/29/10 concentration problems: (show your work for full credit!)

- 1) if you dissolve 20.0 g of NaOH in 250.0 ml of water, what is the molarity of the solution?
- 2) if you dissolve 350.0 g of  $\text{K}_2\text{CrO}_4$  in 2.50 kg of water, what is the molality of the solution?

Concentration worksheet w 328 everett community college student support services program

- 1) 6.80 g of sodium chloride are added to 2750 ml of water. find the mole fraction what is the molarity of a solution in which 0.850 grams of ammonium nitrate are dissolved in 100 ml of water?

Ap chemistry chapter 4. aqueous reactions and solution stoichiometry - 1 - chapter 4. aqueous reactions and solution stoichiometry common student misconceptions

- molarity is moles of solute per liter of solution not per liter of solvent.
- students sometimes used moles instead of molarity in  $M_1V_1 = M_2V_2$

Dilutions worksheet - solutions

- 1) if i add 25 ml of water to 125 ml of a 0.15 m NaOH solution, what will the molarity of the diluted solution be?  
 $M_1V_1 = M_2V_2$  (0.15 m)(125 ml) = x (150 ml) x = 0.125 m
- 2) if i add water to 100 ml of a 0.15 m NaOH solution until the final volume is 150 ml, what will the molarity of the diluted solution be?  
 $M_1V_1 = M_2V_2$

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